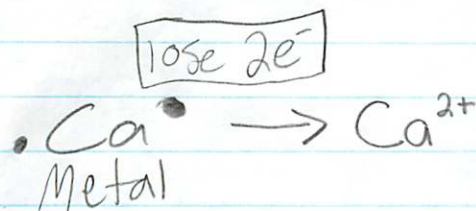
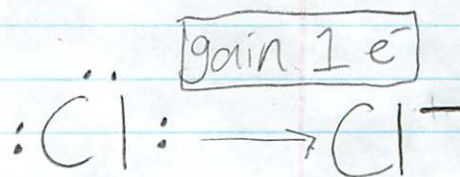


Ch 4 Pt 2 Student Notes

Ion? Ionic compound formation



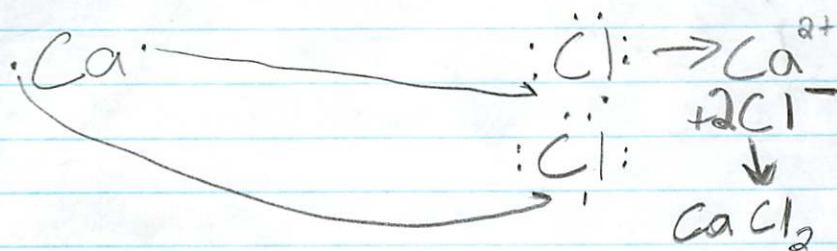
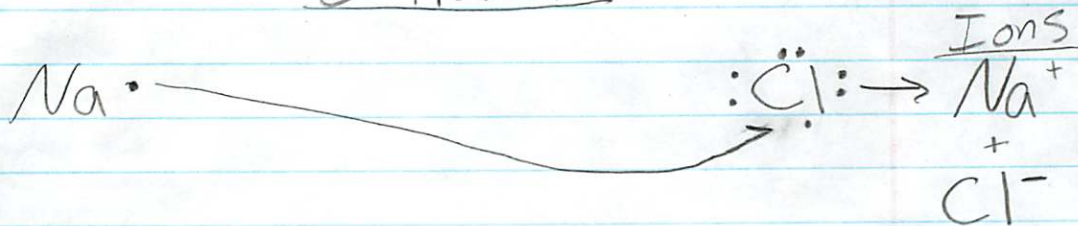
(lose electrons to get octet)



Non-metal
(gain electrons to an octet)

Metals and non-metals combine by exchanging electrons to get a full outer shell. (octet)

e^- Transfer



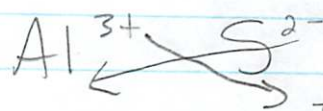
Compound
 NaCl
Formula shows no charges of e^-



Making Formulas for Ionic compounds

	S^{2-}	Br^{-}	NO_3^{-}	SO_4^{2-}	PO_4^{3-}
Na^{+}	Na_2S	$NaBr$	$NaNO_3$	Na_2SO_4	Na_3PO_4
Ca^{+2}	CaS	$CaBr_2$	$Ca(NO_3)_2$	$CaSO_4$	$Ca_3(PO_4)_2$
Al^{+3}	Al_2S_3	$AlBr_3$	$Al(NO_3)_3$	$Al_2(SO_4)_3$	$Al_3(PO_4)_4$

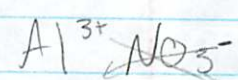
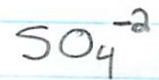
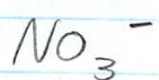
Cross-Met



In an

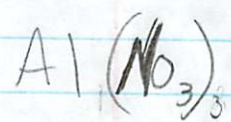
Monatomic Ion - one atom 1 Ion

Polyatomic Ion - Multiple atoms in Ion



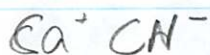
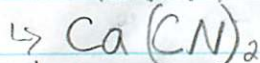
Nitrate

sulfate

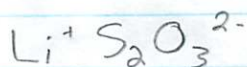


Writing Formulas from Names

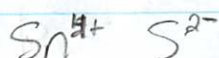
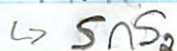
Calcium Cyanide



Lithium Thiosulfate



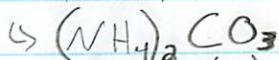
Tin(IV) Sulfide



Sodium Nitrite



Ammonium carbonate



Nickel(III) Chromate

